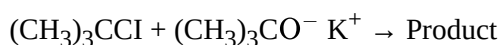


1.



1. S_N Product will be more
2. E_2 Product will be more
3. Both will be the same
4. None of the above

2.

The correct statement regarding electrophile is

1. Electrophile is a negatively charge species and can form a bond by accepting a pair of electrons from a nucleophile
2. Electrophile is a negatively charged species and can form a bond by accepting a pair of electrons from another electrophile
3. Electrophiles are generally neutral species and can form a bond by accepting a pair of electrons from a nucleophile
4. Electrophile can be either neutral or positively charged species and can form a bond by accepting a pair of electrons from a nucleophile

3.

Which of the following statement is not correct for a nucleophile?

1. Nucleophile is a Lewis acid
2. Ammonia is a nucleophile
3. Nucleophiles attack low electrons density sites
4. Nucleophiles are not electron seeking

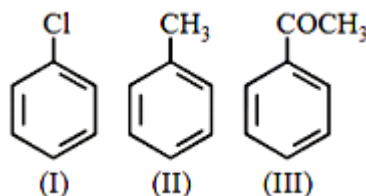
4.

Which of the following compounds will not undergo Friedel-Craft's reaction easily?

1. Cumene
2. Xylene
3. Nitrobenzene
4. Toluene

5.

The increasing order of the reactivity of the following compounds towards electrophilic aromatic substitution reaction is:



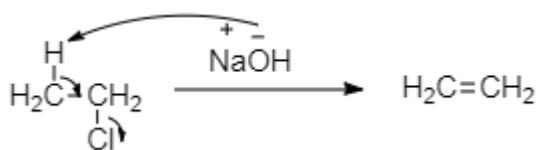
1. $\text{III} < \text{I} < \text{II}$
2. $\text{III} < \text{II} < \text{I}$
3. $\text{II} < \text{I} < \text{III}$
4. $\text{II} < \text{III} < \text{I}$

6.

3-Phenylpropene on reaction with HBr gives (as a major product)-

1. $\text{C}_6\text{H}_5\text{CH}(\text{Br})\text{CH}_2 - \text{CH}_3$
2. $\text{C}_6\text{H}_5\text{CH}_2\text{CH}(\text{Br})\text{CH}_3$
3. $\text{C}_6\text{H}_4(\text{Br})\text{CH}_2\text{CH}_2\text{CH}_3$
4. $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{CH}_2\text{Br}$

7.



Most probable mechanism for this reaction is-

1. E1
2. E2
3. E1_{CB}
4. α elimination

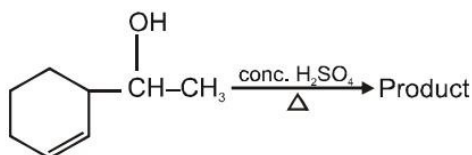
8.

For S_N1 reaction, preferred solvent will be

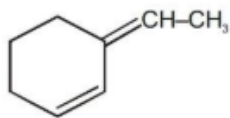
1. Water
2. Benzene
3. Ether
4. Toluene

9.

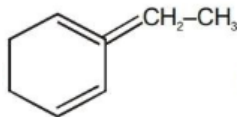
The main product of the following reaction will be :-



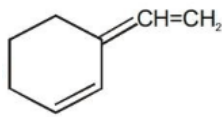
1.



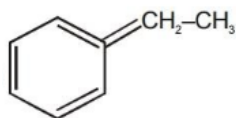
2.



3.



4.



10.

The reaction of propene with HOCl proceeds via the addition of

1. H^+ in the first step
2. Cl^+ in the first step
3. OH^- in the first step
4. Cl^+ and OH^- in a single step

11.

2-Bromopentane reacts with alcoholic KOH to give

1. Cis-2-Pentene
2. Trans-2-Pentene
3. 1-Pentene
4. None of the above

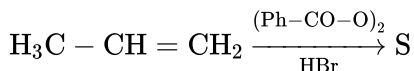
12.

Toluene when refluxed with Br_2 in the presence of light mainly gives

1. p-Bromotoluene
2. Benzyl bromide
3. o-Bromotoluene
4. Mixture of o- and p-bromotoluene

13.

The major product(S) in the below reaction is-



1. $\text{H}_3\text{C} - \text{CH}_2 - \text{CH}_2\text{Br}$
1-Bromopropane

2. $\text{CH}_3 - \underset{\text{Br}}{\text{CH}} - \text{CH}_3$
2-Bromopropane

3. $\text{CH}_3 - \text{CH} = \text{CH}_2$
Propene

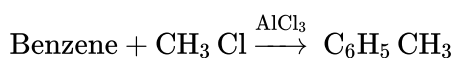
4. None of the above

14.

What type of radicals that can be formed as intermediates during monochlorination of 2-methylpropane ?

1. Primary and tertiary
2. Two types of primary
3. primary and secondary
4. Two types of tertiary

15.

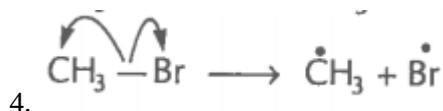
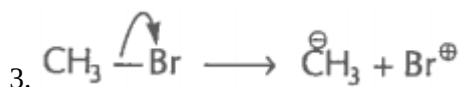
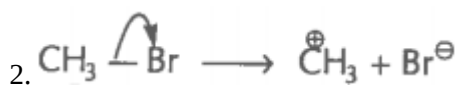
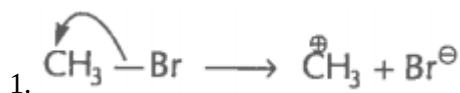


Mechanism & intermediate involve in the above reaction is/are

1. Aromatic electrophilic substitution & carbocation
2. Aromatic Nucleophilic substitution & carbanion
3. Aromatic free radical substitution & Free radical
4. Carbene based substitution reaction & Carbene

16.

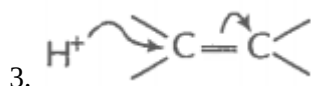
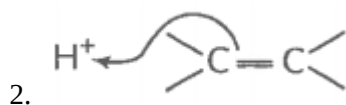
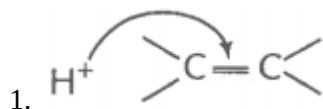
Covalent bond can undergo fission in two different ways. The correct representation involving a heterolytic fission of $\text{CH}_3 - \text{Br}$ is



17.

The addition of HCl to an alkene proceeds in two steps.

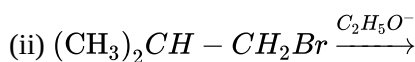
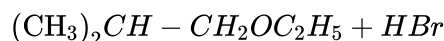
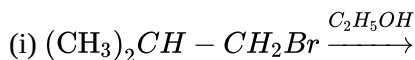
The first step is the attack of H^+ ion to >C=C< portion which can be shown as



4. All of these are possible

18.

Consider the reactions :

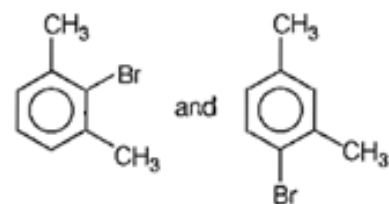
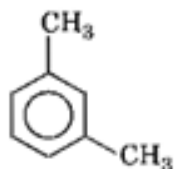


The mechanisms of reactions (i) and (ii) are respectively :

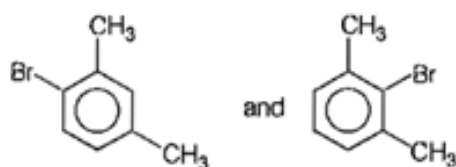
1. $\text{S}_{\text{N}}2$ and $\text{S}_{\text{N}}1$
2. $\text{S}_{\text{N}}1$ and $\text{S}_{\text{N}}2$
3. $\text{S}_{\text{N}}1$ and $\text{S}_{\text{N}}1$
4. $\text{S}_{\text{N}}2$ and $\text{S}_{\text{N}}2$

19.

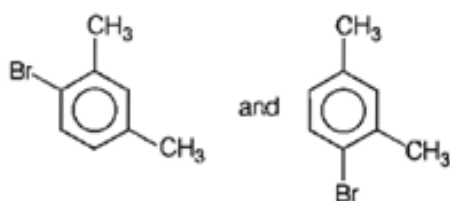
What products are formed when the following compound is treated with Br_2 in the presence of FeBr_3 ?



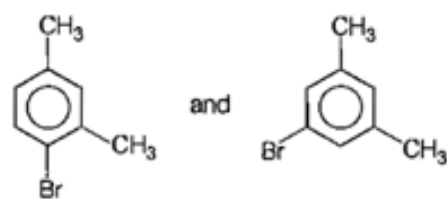
1.



2.



3.



4.

20.

Chlorine atom is a -

1. Carbocation
2. Nucleophile
3. Electrophile
4. Carbanion

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