

1. Transuranium elements are those which are:
 1. Having a higher atomic number than uranium
 2. Lighter than uranium
 3. Having lower atomic number than uranium
 4. Having the same atomic number as uranium

2. Eka aluminium and Eka silicon are now known as:
 1. Ga and Ge
 2. Al and Si
 3. Fe and S
 4. H^+ and Si

3. The lightest metal in the periodic table is:
 1. H
 2. Mg
 3. Ca
 4. Li

4. Which electronic configurations represent to a d block element?
 1. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6$
 2. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^1$
 3. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$
 4. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

5. Elements of group IB and IIB are called:
 1. Normal elements
 2. transition elements
 3. Alkaline earth metals
 4. Alkali metals

6. Elements with electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^3$ belongs to the group of periodic table.
 1. 3rd
 2. 15th
 3. 17th
 4. 12nd

7. The law of triads applicable to a group of:
 1. Cl, Br, I
 2. C, N, O
 3. Na, K, Rb
 4. H, O, N

8. In the long form of periodic table all the non-metals are placed under:
 1. s-block
 2. p-block
 3. f-block
 4. d-block

9. The statements that are not correct for the periodic classification of elements is:
 1. The properties of elements are the periodic functions of their atomic number
 2. Non-metallic elements are lesser in number than metallic elements
 3. The first ionization energies of elements along a period do not vary in a regular manner with an increase in atomic number
 4. For transition elements, the d-subshells are filled with electrons monotonically with the increase in atomic number.

10. Which of the following is not a metal?
 1. Gold
 2. Mercury
 3. Scandium
 4. Selenium

11. The example of metalloid elements in the periodic table is:
1. Na and K
 2. Cu and Al
 3. As and Si
 4. Ca and Mg
12. The element with atomic number 26 is:
1. A non-metal
 2. Krypton
 3. Iron
 4. Manganese
13. The electronic configuration of chalcogens in their outermost shell is,
1. $ns^2 np^3$
 2. $ns^2 np^4$
 3. $ns^2 np^5$
 4. $ns^2 np^6$
14. Electronic configuration of Palladium is :-
1. $[Rn] 5f^3 6d^1 7s^2$
 2. $[Rn] 5f^5 6d^1 7s^2$
 3. $[Rn] 5f^2 6d^1 7s^2$
 4. None of the above
15. An element whose IUPAC name is ununtrium(Uut) belong to:
- (1) s-block element
 - (2) p-block element
 - (3) d-block element
 - (4) Transition element
16. Which of the following is not representative element ?
- (1) Tellurium
 - (2) Tantalum
 - (3) Thallium
 - (4) Astatine
17. Consider the following electronic configuration of an element (P) :
- $$[Xe]4f^{14}5d^16s^2$$
- Then correct statement about element 'P' is :
- (1) It belongs to 6th period and the 1st group
 - (2) It belongs to 6th period and the 2nd group
 - (3) It belongs to 6th period and the 3rd group
 - (4) None of these
18. Which of the following is the incorrect match for atom of element?
1. $[Ar]3d^54s^1 \rightarrow 4^{th} \text{ period, } 6^{th} \text{ group}$
 2. $[Kr]4d^{10} \rightarrow 5^{th} \text{ period, } 12^{th} \text{ group}$
 3. $[Rn]6d^27s^2 \rightarrow 7^{th} \text{ period, } 3^{th} \text{ group}$
 4. $[Xe]4f^{14}5d^26s^2 \rightarrow 6^{th} \text{ period, } 4^{th} \text{ group}$
19. The electronic configurations of Eu (Atomic no. 63), Gd (Atomic no. 64) and Tb (Atomic no. 65) are
1. $[Xe] 4f^6 5d^1 6s^2$, $[Xe]4f^7 5d^1 6s^2$ and $[Xe] 4f^9 6s^2$
 2. $[Xe] 4f^6 5d^1 6s^2$, $[Xe] 4f^7 5d^1 6s^2$ and $[Xe]4f^8 5d^1 6s^2$
 3. $[Xe] 4f^7 6s^2$, $[Xe]4f^7 5d^1 6s^2$ and $[Xe] 4f^9 6s^2$
 4. $[Xe] 4f^7 6s^2$, $(Xe)4f^8 6s^2$ and $[Xe]4f^8 5d^1 6s^2$
20. What type of oxide would Eka- aluminium form?
- (1) EO_4
 - (2) E_2O_3
 - (3) E_3O_2
 - (4) EO

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